BOARD OF INTERMEDIATE EDUCATION, KARACHI

INTERMEDIATE EXAMINATION, 2016 (ANNUAL)

Date: 12-05-2016
9:30 a.m. to 9:50 a.m.

CHEMISTRY PAPER – I

(Science Pre-Engineering & Pre-Medical Groups)

Max. Marks: 17
Time: 20 minutes

The correct answers are highlighted in red colour.

SECTION 'A'
(MULTIPLE CHOICE QUESTIONS) – (M.C.Qs.)

Code No:CH-17

Write this Code No. in the Answerscript.

NOTE:

i) This section consists of 17 part questions and all are to be answered. Each question carries one mark.

ii) Do not copy the part questions in your answerscript. Write only the answer in full against the proper number of the question and its part.

iii) The code of your question paper is to be written in bold letters in the beginning of the answerscript.

1. Choose the correct answer for each from the given options:

i) These have low values of activation energy:
* Slow reactions * **Fast reactions**
* Moderate reactions * Ionic reactions

ii) The bond angle is maximum in this molecule:
* CH_4 * CO_2 * H_2O * NH_3

iii) Gases behave ideally at these conditions:
* High pressures and high temperatures * High pressures and low temperatures
* **Low pressures and high temperatures** * Low pressures and low temperatures

iv) Glass is a/an:
* Crystalline solid * **Amorphous solid** * Covalent solid * Ionic solid

v) Bond Energy is the greatest for:
* CH_4 * O_2 * N_2 * Cl_2

vi) The bulk properties of a system, which are easily measurable, are known as:
* Microscopic properties * Chemical properties
* **Macroscopic properties** * Physical properties

vii) The characteristic of 10^3 is:
* 2 * 3 * 4 * 5

viii) This molecule has zero dipole moment:
* NH_3 * CO_2 * H_2O * HCl

ix) The molarity of a solution containing $20g$ $NaOH$ dissolved into $1dm^3$ solution will be:
* 0.1 * **0.5** * 1 * 2

x) The Octet rule is not valid for this molecule:
* N_2 * CO_2 * O_2 * H_2

xi) The yield of Ammonia in Haber's process is favoured by:
* High pressure and high temperature * **High pressure and low temperature**
* Low pressure and low temperature * Low pressure and high temperature

xii) This ion has greatest degree of hydration:
* Na^+ * Mg^{+2} * Al^{+3} * K^+

xiii) The volume of 3.01×10^{23} molecules of N_2 gas at S.T.P. will be:
* $3dm^3$ * $11.2dm^3$ * $22.4dm^3$ * $28dm^3$

xiv) $Rate = K\ [NH_3]^2$. Keeping the other conditions same, if the concentration of NH_3 is increased by four times, then the initial rate of reaction X will be:
* 2X * 4X * 8X * **16X**

xv) This is not extensive property:
* Entropy * **Viscosity** * Enthalpy * Internal Energy

xvi) The extent of reaction will be maximum for this K_c value:
* 10^{-13} * 0.1 * 10 * 10^3

xvii) This Hydrogen halide has the highest percentage of ionic character:
* **HF** * HCl * HBr * HI

----------xxxxxxxxxx----------