BOARD OF INTERMEDIATE EDUCATION, KARACHI INTERMEDIATE EXAMINATION, 2016 (ANNUAL)

SECTION 'A' MULTIPLE CHOICE OUESTIONS) - (M.C.O.S.) OTTE: i) This section consists of 17 part questions and all are to be answered Fack question carries one mark. ii) Do not copy the part questions and all are to be answered Fack question carries one mark. ii) Do not copy the part questions in your answerscript. Write only the answer in full against the proper number of the question and its part. ii) The code of your question part is to be written in bold letters in the beginning of the answerscript. Interstein code of your question parts is to be written in bold letters in the beginning of the answerscript. Interstein code of your question parts is to be written in bold letters in the beginning of the answerscript. Interstein code of your question part is to be written in bold letters in the beginning of the answerscript. Interstein code of your question parts is to be written in bold letters in the beginning of the answerscript. Interstein code of your question parts is to be written in bold letters in the beginning of the answerscript. Interstein code of your question parts is to be written in bold letters in the beginning of the answerscript. Interstein code of your question part is to be written in bold letters in the beginning of the answerscript. Interstein code of your question parts is be written in bold letters in the beginning of the answerscript. Interstein code of your question parts. Inthis molecules <th answ<="" code="" of="" th=""><th colspan="4">Max. Marks: 17 Time: 20 minutes</th></th>	<th colspan="4">Max. Marks: 17 Time: 20 minutes</th>	Max. Marks: 17 Time: 20 minutes			
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Each question carries one mark. i) Do not copy the part questions in your answerscript. Write only the answer in full against the proper number of the question and its part. ii) The code of your question paper is to be written in bold letters in the beginning of the answerscript. 1. Choose the correct answer for each from the given options: i) These have low values of activation energy: * Slow reactions * Instrument in this molecule: * CH_4 * CO_2 * H_2O * NH_3 iii) The bond angle is maximum in this molecule: * CH_4 * CO_2 * H_2O * NH_3 iii) Gases behave ideally at these conditions: * High pressures and high temperatures * High pressures and low temperatures * Cov pressures and high temperatures * Low pressures and low temperatures iv) Glass is a/an: * $Crystalline solid * Amorphones solid * Covalent solid * Ionic solid v) Bond Energy is the greatest for: * CH_4 * O_2 * N_2 * Cl_2vi) The bulk properties of a system, which are easily measurable, are known as:* Microscopic properties * Dispersives* Macroscopic properties * Coy_3 * H_2O * HClix) The characteristic of 10^3 is:* 2 * O_3 * H_2O * H_2Oiii) This molecule has zero dipole moment:* NH_3 * CO_3 * H_2O * HClix) The molarity of a solution containing 20g NaOH dissolved into 1dm^3 solution will be:* O_1 * O_2 * I * 2x) The Octer rule is not valid for this molecule:* N_2 * CO_2 * I * 2x) The Octer rule is not valid for this molecule:* N_2 * CO_2 * I * 2x) The Victur of this molecule:* N_2 * CO_2 * I* High pressure and low temperature$	cript.				
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xii) This ion has greatest degree of hydration: * Na^+ * Mg^{+2} * Al^{+3} * K^+					
xiii) The volume of 3.01×10^{23} molecules of N_2 gas at S.T.P. will be:					
* $3dm^3$ * $11.2dm^3$ * $22.4dm^3$ * $28dm^3$					
xiv) $Rate = K NH_3^2$. Keeping the other conditions same, if the concentration of NH_3 is increased by four times, then the initial rate of reaction X will be: * $2X$ * $4X$ * $8X$ * 16X					
xv) This is not extensive property: * Entropy * Viscosity * Enthalpy * Internal Energy					
xvi) The extent of reaction will be maximum for this K _c value:					
* 10^{-13} * 0.1 * 10 * 10^{3}					
xvii) This Hydrogen halide has the highest percentage of ionic character: * HF * HCl * HBr * HI					