(MULTAN)

Paper	Code	2	017 (A)		Roll No	
Numbe	er: 2481	INTERMEDIA	TE PART-I (11th CLASS	S)	
TIME A	ALLOWED: 20 M		OBJECTI	<u>VE</u>	MAXIMU	GROUP-I
think is Cutting as given	correct, fill that cit or filling two or m in objective type q	ices for each objective rele in front of that of ore circles will resulutes question paper and le Do not solve question	question numbe It in zero mark i eave others blan	r. Use marke in that question ik. No credit	er or pen to fil on. Attempt as will be award	l the circles. s many question:
(1)	The volume occu	pied by $1.4g$ of N_2	at S.T.P is:-			
	(A) $2.24 dm^3$	(B) 22.4 dm	(C) 1	$12 dm^3$	(D) 112 cm ³	
(2)	The number of Iso	otopes of Cadmium is	-			
	(A) Six	(B) Seven	(C) F	ive	(D) Nine	
(3)	Chromatography i	n which the stationar	y phase is a solic	l is classified a	as:- (A) Partiti	on Chromatograp
	(B) Gas Chromato	graphy (C) Adsor	rption Chromato	graphy (D) Thin layer C	hromatography
(4)	The molar volume	of CO_2 is maximum	at:-			
	(A) S.T.P.	(B) 127°C	and latm (C)	0°C and 2at	m (D) 273°C	and 2 atm
(5)	Pressure remaining	g constant at which te	emperature the v	olume of a gas	will become t	wice of
	what it is at 0°C.		(B) 200°C	(C) 546 K		
(6)	When water freez	es at $0^{\circ}C$, its density	decreases due t	0:-		
		present in the structu			of ice	
	(C) Change of bon	id lengths	(D) C	hang of bond	angles	
(7)	is a pseudos	solid. (A)	CaF_2	(B) Glass	(C) NaCl	(D) <i>KCℓ</i>
(8)	When 6d orbital	is complete, the enter	ring electron goe	es into:-		
	(A) 7 f	(B) 7s	(C) 7	p	(D) 7d	
(9)	The wave number of	of the light emitted by	y a certain source	e is 2 × 10 ⁶ m ⁻	1. The waveler	ngth of this
) 500nm (B) 50			(D) 5×10^7	
(10)		energy in KJ mole i			107404040000000000000000000000000000000	(D) 336
(11)	has zero dip			(GUE) 1/35/6		U411540 1540. 1
nii Es	AND IN THE RESERVE OF THE PARTY		1090-07K - 100 100 100 100 100 100 100 100 100 1	(B) CHCl ₃		(D) BF ₃
	(A) Enthalpy change	energy of a chemical				
		ole dm ⁻³ of an aqueo	sublimation (C)	31	(D) Internal	energy change
	(A) 3.0	(B) 2.7			(T)\ 1.6	
(14)		by C.M. Guldberg and	(C) 2. d P. Waane in 18		(D) 1.5	ation of Mass
	(B) Law of Mass Ac		oution Law			ation of Energy
		olved in 90 g of water				
	(A) $\frac{1}{5}$	(B) 5.1	(C) $\frac{1}{51}$		(D) 6	700 - 100 - 100 O
(16)	The potential of sta	ndard Hydrogen Elec	trode is arbitrari	ly taken as:-		
	(A) 1.00	(B) 0.00	(C) 5.0	00	(D) 3.00	
		d into ethanol by the	004104 (CT)(15100-5611)			
	(A) Invertase	(B) Urease	(C) Zv	mase	(D) Sucrase	

21(Obj) (\$\frac{1}{2}\$)-2017(A)-18000 (MULTAN)

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Paper	Code

Roll	No.			
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2483 INTERMEDIATE PART-I (11th CLASS)

	MISTRY PAPER-I			SION 2015-2017)	GROUP-I
	ALLOWED: 20 Minut		BJECTIVE		UM MARKS: 17
	You have four choices fo s correct, fill that circle in				
uttin	g or filling two or more ci	rcles will result in z	ero mark in that	question. Attempt a	s many questions
s give	n in objective type question LES are not filled. Do no	on paper and leave of t solve question on	others blank. No this sheet of OB	o credit will be award JECTIVE PAPER.	ded in case
).No.1					
(1)	The molar volume of Co	OWNER CHARACTERS CONTRACTOR			
	(A) S.T.P.			and 2 atm (D) 273° (
(2)	Pressure remaining cons		ature the volume	of a gas will become	twice of
	what it is at $0^{\circ}C$. (A) 546°C (B) 200°C (C) 546 K (D) 273 K	
(3)	When water freezes at ($0^{\circ}C$, its density decr	eases due to:-		
	(A) Empty spaces preser	nt in the structure of	ice (B) Cubic st	ructure of ice	
	(C) Change of bond leng	ths	(D) Chang o	f bond angles	
(4)	is a pseudosolid.	(A) CaF ₂	(B) (Glass (C) NaCl	(D) <i>KC</i> ℓ
(5)	When $6d$ orbital is con	nplete, the entering e	lectron goes into:		
	(A) 7f	(B) 7s	(C) 7 p	(D) 7d	
(6)	The wave number of the	light emitted by a cer	rtain source is 2	< 10 ⁶ m ⁻¹ . The waveler	ngth of this
	light will be:- (A) 500		(C) 200 nm		383 - 3 CANADA
(7)	The $H-H$ bond energy	in KJ mole" is:-	(A) 346 (B) 436 (C) 463	(D) 336
(8)	has zero dipole mo	ment. (A) NH ₃ (B)	CHCℓ₃ (C) H₂O	(D) BF.
(9)	The change in heat energy				
	(A) Enthalpy change			energy (D) Internal	
(10)	The pH of 10^{-3} mole dm				Syange
	(A) 3.0	(B) 2.7	(C) 2.0	(D) 1.5	
(11)	was derived by C.N	1. Guldberg and P. W		2. 6	ation of Mass
	(B) Law of Mass Action	(C) Distribution		(D) Law of Conserva	
(12)	18g glucose is dissolved	in 90 g of water. Th	e relative lowerin		
	(A) $\frac{1}{5}$	(B) 5.1	(C) $\frac{1}{51}$	(D) 6	
(13)	The potential of standard	Hydrogen Electrode	is arbitrarily take	n as:-	
	(A) 1.00	(B) 0.00	(C) 5.00	(D) 3.00	
(14)	Glucose is converted into	ethanol by the enzym	ie:-		
	(A) Invertase	(B) Urease	(C) Zymase	(D) Sucrase	
(15)	The volume occupied by	$1.4 g$ of N_2 at S.T.I	P is:-		
	(A) 2.24 dm ³	(B) $22.4 dm^3$	(C) $1.12 dm^3$	(D) 112 cm ³	
(16)	The number of Isotopes o	f Cadmium is:-			
	(A) Six	(B) Seven	(C) Five	(D) Nine	
(17)	Chromatography in which		is a solid is class	ified as:- (A) Partitio	n Chromatography
	(B) Gas Chromatography	(C) Adsorption (Chromatography	(D) Thin layer Ch	romatography

21(Obj) (\$\frac{1}{2} \frac{1}{2} \)-2017(A)-18000 (MULTAN)

Paper Code Number: 2485 INTERMEDIATE PART-I (11th CLASS) CHEMISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017) GR TIME ALLOWED: 20 Minutes OBJECTIVE MAXIMUM M Note: You have four choices for each objective type question as A, B, C and D. The choice whithink is correct, fill that circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of the pen and leave others blank. No credit will be awarded in cell fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question number. Use marker or pen to fill the circle in front of that question num	OUP-I ARKS: 17 ich you rcles. questions ase ed:- change Mass Energy
CHEMISTRY PAPER-I (NEW SCHEME) (SESSION 2015-2017) GR TIME ALLOWED: 20 Minutes OBJECTIVE MAXIMUM M Note: You have four choices for each objective type question as A, B, C and D. The choice with think is correct, fill that circle in front of that question number. Use marker or pen to fill the circuiting or filling two or more circles will result in zero mark in that question. Attempt as many as given in objective type question apper and leave others blank. No credit will be awarded in c BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER. Q.No.1 (A) Enthalpy change (B) Heat of sublimation (C) Bond energy (D) Internal energy (A) 3.0 (B) 2.7 (C) 2.0 (D) 1.5 (A) 3.0 (B) 2.7 (C) 2.0 (D) 1.5 (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) Law of Mass Action (A) $\frac{1}{5}$ (B) 5.1 (C) $\frac{1}{51}$ (D) 6 (D) 6 (D) 6 (D) 6 (E) Conservation of (C) Soon (D) 3.00 (D) 3.00 (E) Conservation of (C) Soon (D) Socrase (E) Conservation of (C) Soon (D) Socrase	ARKS: 17 ich you rcles. questions ase ed:- change Mass Energy
TIME ALLOWED: 20 Minutes OBJECTIVE MAXIMUM M. Note: You have four choices for each objective type question as A, B, C and D. The choice whithink is correct, fill that circle in front of that question number. Use marker or pen to fill the circle tricles will result in zero mark in that question. Attempt as many as given in objective type question paper and leave others blank. No credit will be awarded in c BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER. Q.No.1 (A) Enthalpy change (B) Heat of sublimation (C) Bond energy (D) Internal energy (A) Enthalpy change (B) Heat of sublimation (C) Bond energy (D) Internal energy (D)	ARKS: 17 ich you rcles. questions ase ed:- change Mass Energy
think is correct, fill that circle in front of that question number. Use marker or pen to fill the circutting or filling two or more circles will result in zero mark in that question. Attempt as many as given in objective type question paper and leave others blank. No credit will be awarded in c BUBBLES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER. Q.No.1 (1) The change in heat energy of a chemical reaction at constant temperature and pressure is called (A) Enthalpy change (B) Heat of sublimation (C) Bond energy (D) Internal energy	ed:- change Mass Energy
(1) The change in heat energy of a chemical reaction at constant temperature and pressure is calle (A) Enthalpy change (B) Heat of sublimation (C) Bond energy (D) Internal energy (2) The pH of 10 ⁻³ mole dm ⁻³ of an aqueous solution of H ₂ SO ₄ is:- (A) 3.0 (B) 2.7 (C) 2.0 (D) 1.5 (3) was derived by C.M. Guldberg and P. Waage in 1864. (A) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (D) Law	Mass Energy
(2) The pH of 10 ⁻³ mole dm ⁻³ of an aqueous solution of H_2SO_4 is:- (A) 3.0 (B) 2.7 (C) 2.0 (D) 1.5 (3) was derived by C.M. Guldberg and P. Waage in 1864. (A) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) 18 g glucose is dissolved in 90 g of water. The relative lowering of vapour pressure is equal (A) $\frac{1}{5}$ (B) 5.1 (C) $\frac{1}{51}$ (D) 6 (5) The potential of standard Hydrogen Electrode is arbitrarily taken as:- (A) 1.00 (B) 0.00 (C) 5.00 (D) 3.00 (6) Glucose is converted into ethanol by the enzyme:- (A) Invertase (B) Urease (C) Zymase (D) Sucrase (7) The volume occupied by 1.4 g of N_2 at S.T.P is:- (A) 2.24 dm ³ (B) 22.4 dm ³ (C) 1.12 dm ³ (D) 112 cm ³ (8) The number of Isotopes of Cadmium is:- (A) Six (B) Seven (C) Five (D) Nine (9) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography The molar volume of CO_2 is maximum at:-	Mass Energy
(A) 3.0 (B) 2.7 (C) 2.0 (D) 1.5 — was derived by C.M. Guldberg and P. Waage in 1864. (A) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (D) 18 g glucose is dissolved in 90 g of water. The relative lowering of vapour pressure is equal (A) $\frac{1}{5}$ (B) 5.1 (C) $\frac{1}{51}$ (D) 6 (5) The potential of standard Hydrogen Electrode is arbitrarily taken as:- (A) 1.00 (B) 0.00 (C) 5.00 (D) 3.00 (6) Glucose is converted into ethanol by the enzyme:- (A) Invertase (B) Urease (C) Zymase (D) Sucrase (7) The volume occupied by 1.4 g of N_2 at S.T.P is:- (A) 2.24 dm³ (B) 22.4 dm³ (C) 1.12 dm³ (D) 112 cm³ (8) The number of Isotopes of Cadmium is:- (A) Six (B) Seven (C) Five (D) Nine (9) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (10) The molar volume of CO_2 is maximum at:-	Energy
(a) was derived by C.M. Guldberg and P. Waage in 1864. (A) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of 18 g glucose is dissolved in 90 g of water. The relative lowering of vapour pressure is equal (A) \frac{1}{5} \qquad (B) 5.1 \qquad (C) \frac{1}{51} \qquad (D) 6 (5) The potential of standard Hydrogen Electrode is arbitrarily taken as:- (A) 1.00 (B) 0.00 (C) 5.00 (D) 3.00 (6) Glucose is converted into ethanol by the enzyme:- (A) Invertase (B) Urease (C) Zymase (D) Sucrase (7) The volume occupied by 1.4 g of N ₂ at S.T.P is:- (A) 2.24 dm ³ (B) 22.4 dm ³ (C) 1.12 dm ³ (D) 112 cm ³ (8) The number of Isotopes of Cadmium is:- (A) Six (B) Seven (C) Five (D) Nine (9) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (10) The molar volume of \$CO_2\$ is maximum at:-	Energy
(3) was derived by C.M. Guldberg and P. Waage in 1864. (A) Law of Conservation of (B) Law of Mass Action (C) Distribution Law (D) Law of Conservation of (4) 18 g glucose is dissolved in 90 g of water. The relative lowering of vapour pressure is equal (A) \frac{1}{5} (B) 5.1 (C) \frac{1}{51} (D) 6 (5) The potential of standard Hydrogen Electrode is arbitrarily taken as:- (A) 1.00 (B) 0.00 (C) 5.00 (D) 3.00 (6) Glucose is converted into ethanol by the enzyme:- (A) Invertase (B) Urease (C) Zymase (D) Sucrase (7) The volume occupied by 1.4 g of \(N_2 \) at S.T.P is:- (A) 2.24 dm³ (B) 22.4 dm³ (C) 1.12 dm³ (D) 112 cm³ (8) The number of Isotopes of Cadmium is:- (A) Six (B) Seven (C) Five (D) Nine (9) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chro (B) Gas Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (10) The molar volume of \(CO_2 \) is maximum at:-	Energy
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(5) The potential of standard Hydrogen Electrode is arbitrarily taken as: (A) 1.00 (B) 0.00 (C) 5.00 (D) 3.00 (E) 3.00 (E) 3.00 (E) 3.00 (D) 3.00 (E) 3.00	
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 (A) 1.00 (B) 0.00 (C) 5.00 (D) 3.00 (E) Glucose is converted into ethanol by the enzyme:- (A) Invertase (B) Urease (C) Zymase (D) Sucrase (The volume occupied by 1.4 g of N₂ at S.T.P is:- (A) 2.24 dm³ (B) 22.4 dm³ (C) 1.12 dm³ (D) 112 cm³ (B) The number of Isotopes of Cadmium is:- (A) Six (B) Seven (C) Five (D) Nine (P) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chromatography (B) Gas Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (10) The molar volume of CO₂ is maximum at:- 	
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(7) The volume occupied by 1.4 g of N ₂ at S.T.P is:- (A) 2.24 dm ³ (B) 22.4 dm ³ (C) 1.12 dm ³ (D) 112 cm ³ (8) The number of Isotopes of Cadmium is:- (A) Six (B) Seven (C) Five (D) Nine (P) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chromatography (B) Gas Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (D) The molar volume of CO ₂ is maximum at:-	
(A) 2.24 dm³ (B) 22.4 dm³ (C) 1.12 dm³ (D) 112 cm³ (B) The number of Isotopes of Cadmium is:- (A) Six (B) Seven (C) Five (D) Nine (C) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chromatography (B) Gas Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (D) The molar volume of CO₂ is maximum at:-	
(8) The number of Isotopes of Cadmium is:- (A) Six (B) Seven (C) Five (D) Nine (9) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chro (B) Gas Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (10) The molar volume of CO ₂ is maximum at:-	
(A) Six (B) Seven (C) Five (D) Nine (P) Chromatography in which the stationary phase is a solid is classified as:- (A) Partition Chromatography (B) Gas Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (D) The molar volume of CO ₂ is maximum at:-	
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(B) Gas Chromatography (C) Adsorption Chromatography (D) Thin layer Chromatography (10) The molar volume of CO_2 is maximum at:-	
(10) The molar volume of CO_2 is maximum at:-	matography
	graphy
(A) S.T.P. (B) 127°C and 1 atm (C) 0°C and 2 atm (D) 273°C and 2.	
(5) v = and 2 and (b) 2/3 C and 2 to	<i>atm</i>
(11) Pressure remaining constant at which temperature the volume of a gas will become twice of	
what it is at $0^{\circ}C$. (A) $546^{\circ}C$ (B) $200^{\circ}C$ (C) $546K$ (D) $273K$	
(12) When water freezes at $0^{\circ}C$, its density decreases due to:-	
(A) Empty spaces present in the structure of ice (B) Cubic structure of ice	
(C) Change of bond lengths (D) Chang of bond angles	
(13) is a pseudosolid. (A) CaF_2 (B) Glass (C) $NaC\ell$ (D) H	KCl
(14) When 6d orbital is complete, the entering electron goes into:-	CALCARD .
(A) $7f$ (B) $7s$ (C) $7p$ (D) $7d$	
(15) The wave number of the light emitted by a certain source is $2 \times 10^6 m^{-1}$. The wavelength of the	is
light will be:- (A) $500 nm$ (B) $500 m$ (C) $200 nm$ (D) $5 \times 10^7 m$	911
(16) The $H-H$ bond energy in $KJ mole^{-1}$ is:- (A) 346 (B) 436 (C) 463 (D) 3	36
(17) has zero dipole moment. (A) NH_3 (B) $CHC\ell_3$ (C) H_2O (D) E	

21(Obj) (\$ \$ \$ \$ \$ \$ \$)-2017(A)-18000 (MULTAN)

Paper	Code	2	017 (A)		Roll No	
Numbe	er: 2487	INTERMEDIA	TE PART-I (11 th CLAS	SS)	
	No.					GROUP-
	ALLOWED: 20	ER-I (NEW SCE Minutes	OBJECTI		MAXIMU	JM MARKS
Note: think is Cutting as given	You have four cho correct, fill that ci or filling two or n in objective type	oices for each objective in front of that of the core circles will result question paper and lead to not solve question.	e type question question numbe t in zero mark i eave others blar	as A, B, C : r. Use mark in that ques ik. No cred	ker or pen to fill tion. Attempt as it will be award	the circles. many questi
(1)	was derived	by C.M. Guldberg an	d P. Waage in 18	864. (A)	Law of Conserva	ation of Mass
	(B) Law of Mass A	Action (C) Distri	bution Law	(D)	Law of Conserv	ation of Energ
(2)	18 g glucose is dis	ssolved in 90 g of wat	er. The relative	lowering of	vapour pressure	is equal to:-
	(A) $\frac{1}{5}$	(B) 5.1	(C) $\frac{1}{5}$	1	(D) 6	
(3)	The potential of s	tandard Hydrogen Ele	ctrode is arbitrar	ily taken as:	•	
	(A) 1.00	(B) 0.00	(C) 5.	00	(D) 3.00	
(4)	Glucose is conver	ted into ethanol by the	enzyme:-			
	(A) Invertase	(B) Urease	(C) Z	ymase	(D) Sucrase	
(5)	The volume occu	upied by 1.4 g of N_2	at S.T.P is:-			
	(A) $2.24 dm^3$	(B) 22.4 dm	r³ (C) 1.	12 <i>dm</i> ³	(D) $112cm^3$	
(6)	The number of Is	otopes of Cadmium is	:4			
	(A) Six	(B) Seven	(C) Fi	ive	(D) Nine	
(7)	Chromatography	in which the stationar	y phase is a solid	l is classified	l as:- (A) Partitio	on Chromatog
	(B) Gas Chromat		rption Chromato	graphy	(D) Thin layer Cl	hromatograph
(8)	The molar volum	e of CO ₂ is maximum	n at:-			
	(A) S.T.P.	(B) 127°C	and 1 atm (C)	0"C and 2	atm (D) 273°C	and 2 atm
(9)	Pressure remaining	ng constant at which to	emperature the vo	olume of a g	as will become t	wice of
	what it is at $0^{\circ}C$. (A) 546°C	(B) 200°C	(C) 546	K (D) 273 K	
(10)	When water fre	ezes at 0°C, its densi	ty decreases due	to:-		
	(A) Empty spaces	s present in the structu	re of ice (B) C	ubic structur	e of ice	
	(C) Change of bo	and lengths	(D) C	hang of bone	d angles	
(11)	is a pseud	osolid. (A)	CaF_2	(B) Glass	(C) NaCℓ	(D) <i>KCℓ</i>
(12)	When 6d orbital	is complete, the enter	ing electron goes	s into:-		
	(A) 7f	(B) 7s	(C) 7	p	(D) 7d	
(13)	The wave number	of the light emitted by	y a certain source	e is 2 × 10 ⁶ r	n ⁻¹ . The waveler	igth of this
	light will be:- (A	A) 500 nm (B) 50	00 m (C) 20	00 n m	(D) 5×10^7	m
(14)	The $H-H$ bond	energy in KJ mole-1	is:- (A) 346	(B) 436	(C) 463	(D) 336
(15)		pole moment.	(A) NH ₃	(3) (6)	(C) H ₂ O	(D) BF ₃
		energy of a chemical	S S S	2.3		
W. S. T. W.	(A) Enthalpy chan			V. 10.	gy (D) Internal	
		· V	STATE OF THE STATE		- 1	O/

(A) 3.0 (B) 2.7 (C) 2.0

Roll No:

INTERMEDIATE PART-I (11th CLASS) PAPER-I (NEW SCHEME) CHEMISTRY (SESSION 2015-2017) GROUP-II TIME ALLOWED: 2.40 Hours SUBJECTIVE MAXIMUM MARKS: 68 NOTE: - Write same question number and its part number on answer book, as given in the question paper. SECTION-I 2. Attempt any eight parts. $8 \times 2 = 16$ N_2 and CO have same numbers of Electrons, Protons and Neutrons. Justify it, with reason. (i) Define Molecular Formula. Give two examples of the compounds having same empirical (ii) and molecular formulas. No individual atom of Neon in the sample has a mass of 20.18 a.m.u. Give reason. (iii) (iv) Define Crystallization. What is basic principle of crystallization? What is the difference between Adsorption and Partition Chromatography? (v) Write two salient features of an ideal solvent used in the process of Crystallization. (vi) Calculate the value of Ideal gas constant 'R' in S.I. units. (vii) What is Joule Thomson Effect? (viii) Write two applications of Plasma. (ix) (x) Give statement of "Law of Mass Action". What is the effect of Catalyst on equilibrium position of a reaction? (xi) (xii) Explain that a mixture of NH_4OH and $NH_4C\ell$ gives us the basic buffer. 3. Attempt any eight parts. $8 \times 2 = 16$ Why the boiling point of water is different at Murree hills and Mount Everest? (i) The values of boiling points of noble gases increase from top to bottom within a group. Give reason. (ii) Define Unit Cell. Give one example. (iii) The electrical conductivity of metals decreases with increase in temperature. Why? (iv) Sate Hund's rule. Give one example. (v) (vi) What is meant by fine structure of Hydrogen Spectrum? (vii) What are X – rays? What is their origin? Write balanced equations for any two nuclear reactions. (viii) What is the difference between Zeotropic and Azeotropic solutions? (ix) (x) What are Discontinuous Solubility Curves? Give one example. (xi) What is Anodized Aluminium? Write redox reactions which occur during discharging of lead accumulator battery? (xii) Attempt any six parts. $6 \times 2 = 12$ (i) The size of a cation is smaller than its parent atom. Prove. Define Ionization Energy (IE) and Electron Affinity(EA). (ii) (iii) The dipole moments of CO_2 and CS_2 are zero, but that of SO_2 is 1.61D. Give reasons. Why bond formation is not possible between two He atoms. (iv) Prove with Molecular Orbital Theory(MOT)? State the first Law of Thermochemistry. (v) Burning of a candle is a spontaneous process. Give reason. (vi) A finally divided catalyst may prove more effective. Give reason. (vii) Write two examples of Enzyme Catalyzed reactions. (viii) What is Pseudo First Order Reaction? (ix) SECTION-II NOTE: - Attempt any three questions. A sample of liquid consisting of Carbon, Hydrogen and Oxygen was subjected to combustion 5.(a) analysis. 0.5439 g of compound gave 1.039 g of CO2, 0.6369 g of water. Determine the empirical formula of the compound. What are Ionic Solids? Write six properties of Ionic Solids. (b) 4 State and explain Boyle's Law and verify this Law by an experiment. 6.(a) 4 What is Cathode Ray Tube? Describe two properties of Cathode Rays. (b) Describe Valence Shell Electron pair Repulsion Theory and give its postulates. 7.(a) Give example of structure of Ammonia Molecule by this theory. Define Enthalpy and derive Enthalpy change at constant pressure. (b) What is Voltaic Cell? Explain with one example. 8.(a) The solubility of PbF_2 at $25^{\circ}C$ is $0.64 \, gm/dm^3$. Calculate K_{SP} of PbF_2 . (b) 4 Explain Lowering of Vapour Pressure by adding a Non-volatile, Non electrolyte solute 9.(a) in a solvent.

What is Half Life Period? Give examples, also give its mathematical form.

22-2017(A)-12000

(MULTAN)

4

(b)

Paper Code	2017 (A)	Roll No	

Number: 248

2482 INTERMEDIATE PART-I (11th CLASS)

Numb	ber: 2-TOZ	I EIGHE DINTE I MIN	T(II CENSS)	
CHE	MISTRY PAPER-I	(NEW SCHEME)	(SESSION 2015-	2017) GROUP-II
ГІМЕ	ALLOWED: 20 Minu	tes <u>OBJEC</u>	CTIVE M	AXIMUM MARKS: 17
think is Cutting as give	s correct, fill that circle in g or filling two or more c n in objective type questi LES are not filled. Do n	or each objective type quest in front of that question nu- ircles will result in zero ma on paper and leave others ot solve question on this sh	mber. Use marker or p ark in that question. At blank. No credit will b	oen to fill the circles. tempt as many questions be awarded in case
(1)	The mass of one mole	of electron is:-	4	
	(A) 1.008 g	(B) 0.55mg	(C) 0.184 mg	(D) 1.673 mg
(2)	The largest number of r	nolecules are present in:-		
	(A) $3.6 g$ of H_2O	(B) $4.8g$ of C_2H_5OH	(C) 2.8 g of CO	(D) 5.4 g of N_2O_5
(3)	Solvent extraction is an	equilibrium process and it i	s controlled by:- (A)	Law of Mass Action
	(B) The amount of Solv	ent used (C) Distribution	on Law (D) The	amount of Solute
(4)	If absolute temperature	of a gas is doubled and pres	sure is reduced to one ha	alf, the volume of gas will:-
	(A) Remain unchanged	(B) Increase four times	(C) Reduce to $\frac{1}{4}$	(D) Be doubled
(5)	The order of rate of diff	fusion of gases NH_3 , SO_2 ,	$C\ell_2$ and CO_2 is:- (A)	$NH_3 > SO_2 > C\ell_2 > CO_2$
	(B) $NH_3 > CO_2 > SO_2$	$> C\ell_2$ (C) $C\ell_2 > SO_2$	$> CO_2 > NH_3$ (D)	$NH_3 > CO_2 > C\ell_2 > SO_2$
(6)		re water at 1 atm pressure		
	(A) 98°C	(B) 100° C	(C) 69°C	(D) 120°C
(7)	Ionic solids are characte		West State of the	(D) 120 C
327 3487		n solid state (C) High vap		ility in polar solvents
(8)		aplete, the entering electron		• •
	(A) 7f	(B) 7s	(C) 7p	(D) 7d
(9)	The velocity of Photon	is:- (A) Independent	7.07	(2)
	(B) Equal to square of its	s amplitude (C) Depends) Depends on its source
(10)		has the highest percentage of		
	(A) HCl	(B) <i>HBr</i>	(C) HF	(D) HI
(11)	The bond order of N_2 is:	- (A) 2	(B) 3 (C) 4	(D) 1
(12)	For the reaction NaOH	$+$ $HC\ell \longrightarrow NaC\ell + H_2O$	the change in enthalpy	is called:-
) Heat of formation (C) H		
(13)		of an aqueous solution of		,
	(A) 3.0	(B) 2.7	(C) 2.0	(D) 1.5
(14)	The ionization constant of	of pure water at 25°C is:-	14	15/8
		(C) 1.0 × 10 ⁻¹⁴ mole ²		
(15)		constant is the ratio of the ele	67.08	
100000	(A) Molarity	(B) Molality (C) Mole		
(16)		Fluorine in OF_2 is:-		
				(-)

(17) The unit of the rate constant is the same as that of rate of reaction in:-

(B) 2nd order reaction

(A) 1st order reaction

22(Obj) ()-2017(A)-12000 (MULTAN)

(C) Zero order reaction (D) 3rd order reaction

Paper	Code		2017 (A)		Roll No	
Numbe	er: 2484	INTERMEI	DIATE PART	-I (11th CLAS	SS)	
CHEN	IISTRY PAPI	ER-I (NEW S	SCHEME)	(SESSION	2015-2017)	GROUP-II
TIME .	ALLOWED: 20 1	Minutes	OBJEC	CTIVE	MAXIMU	JM MARKS: 17
think is Cutting as giver	You have four choice correct, fill that city or filling two or many in objective type of ES are not filled.	rcle in front of theore circles will requestion paper a	hat question nu esult in zero m nd leave others	mber. Use marl ark in that quest blank. No cred	ker or pen to fil tion. Attempt as it will be award	l the circles. s many questions
(1)	If absolute temper	rature of a gas is	doubled and pres	sure is reduced to	o one half, the v	olume of gas will:
	(A) Remain uncha	anged (B) Inc	rease four times	(C) Reduc	e to $\frac{1}{4}$ (D)	Be doubled
(2)	The order of rate	of diffusion of ga	ses NH_3 , SO_2 ,	$C\ell_2$ and CO_2	s:- (A) NH ₁ >	$SO_{\gamma} > C\ell_{\gamma} > CO$
	(B) $NH_3 > CO_2$	$> SO_2 > C\ell_2$	(C) $C\ell_2 > SO_2$	$> CO_2 > NH_3$	(D) NH ₃ >	$CO_2 > C\ell_2 > SO_2$
(3)	The boiling point	of pure water at	l atm pressure	is:-		
	(A) 98°C	(B) 100	$O^{\alpha}C$	(C) 69° C	(D)	120°C
(4)	Ionic solids are cl	aracterized by:-	(A) Low n	nelting points		
	(B) Good conduct	tivity in solid stat	e (C) High var	our pressure (D) Solubility in p	olar solvents
(5)	When 6d orbital	is complete, the	entering electror	goes into:-		
	(A) 7f	(B) 7s		(C) 7p	(D)	7 <i>d</i>
(6)	The velocity of Pl	hoton is:-	(A) Independent	of its wavelengt	h	
	(B) Equal to squar	e of its amplitude	(C) Depend	s on its waveleng	gth (D) Deper	nds on its source
(7)		alide has the high	nest percentage of	of ionic character		
	(A) <i>HCℓ</i>	(B) HE		(C) HF	(D)	HI
(8)	The bond order of	N ₂ is:-	(A) 2	(B) 3	(C) 4	(D) 1
(9)	For the reaction A	NaOH + HCl —	$\rightarrow NaC\ell + H_2$	O the change in	enthalpy is calle	d:-
	(A) Heat of reaction	on (B) Heat of f	ormation (C)	Heat of Neutraliz	ation (D) Heat	of Combustion
(10)	The pH of 10^{-3} me	ole dm ⁻³ of an aq	ueous solution o	f <i>H</i> ₂ <i>SO</i> ₄ is:-		
	(A) 3 0	(B) 2.7	(2)	(C) 2.0	(D)	1.5

(11) The ionization constant of pure water at 25° C is:-

(13) The oxidation number of Fluorine in OF₂ is:-

(16) The largest number of molecules are present in:-

(15) The mass of one mole of electron is:-

(A) 1st order reaction (B) 2nd order reaction

(A) Molarity

(A) 1.008 g

(A) 3.6 g of H_2O

(B) 1.6×10^{-14} mole dm^{-3} (C) 1.0×10^{-14} mole 2 dm^{-6}

(12) The molal boiling point constant is the ratio of the elevation in boiling point to:-

The unit of the rate constant is the same as that of rate of reaction in:-

(B) 0.55 mg

(B) The amount of Solvent used (C) Distribution Law

(A) 1.8×10^{-16} mole dm⁻³

(A) + 2 (B) - 2 (C) - 1

(B) Molality (C) Mole fraction of solute (D) Mole fraction of solvent

(C) 0.184 mg

(B) 4.8 g of $C_2 H_5 OH$ (C) 2.8 g of CO

Solvent extraction is an equilibrium process and it is controlled by:- (A) Law of Mass Action

(D) 1.8 × 10⁻¹⁴ mole² dm⁻⁶

(C) Zero order reaction (D) 3rd order reaction

(D) The amount of Solute

22(Obj) (\$\frac{1}{2} \frac{1}{2} \rightarrow 2017(A)-12000 (MULTAN)

(D) 1.673 mg

(D) 5.4 g of N_2O_5

(D) + 1

Paper	Code		20)17 (A)		Roll No.	
Numb	er: 24	86 INT	TERMEDIA"	ΓΕ PART-I	(11th CLAS		
			(MENNI OCH	ED S. A. SED.	(CECCION	*****	Spare -
			(NEW SCH			2015-2017)	GROUP-II
TIME	ALLOWEI	D: 20 Minut	es	OBJECT	IVE	MAXIMU	JM MARKS: 17
think is Cutting as given	correct, fill or filling two in objective ES are not f	that circle in o or more cir type question	front of that q	uestion numb in zero mark ave others bla	er. Use mark in that quest ank. No credi	t will be award	the circles.
(1)		der of N_2 is:		(A) 2	(B) 3	(C) 4	(D) 1
(2)	For the react	ion NaOH -	$+$ $HC\ell \longrightarrow N\ell$	$aC\ell + H_2O$ th	he change in e	nthalpy is called	<u>.</u>
	(A) Heat of	reaction (B	Heat of format	ion (C) Hea	t of Neutraliza	ation (D) Heat	of Combustion
(3)		CHESCO DE PERSON	of an aqueous			8 8	
	(A) 3.0		(B) 2.7		(C) 2.0	(D)	1.5
(4)	The ionization	on constant of	pure water at 2	25" C is:-	(A) 1.8 × 1	3.76	
	(B) 1.6 × 10	-14 mole dm-1	(C) 1.0 ×	$10^{-14} \text{ mole}^2 ds$	m^{-6} ((D) $1.8 \times 10^{-14} n$	nole2 dm-6
(5)	The molal b	oiling point c	onstant is the ra	tio of the elev	ation in boiling	g point to:-	
(6)	(A) Molarit						ction of solvent
(6)			Fluorine in OF) -2 (C) -	- I (D) + 1
(7)	(A) 1 st order		ant is the same a	reaction			v ard
(8)		f one mole of		reaction	(C) Zero or	der reaction (D) 3 rd order reaction
11,745,755	(A) 1.008 g		(B) 0.55 mg		(C) 0.184 m	(D) I	.673 mg
(9)	The largest	number of me	olecules are pre-	sent in:-		6 (D) 1	iorsing
					(C) 2.8 g of	* CO (D) 5	Ag of NO
(10)						(A) Law of N	
		ount of Solve		Distribution L) The amount o	
(11)	If absolute t	emperature o			200		lume of gas will:-
		unchanged				to $\frac{1}{4}$ (D) B	225
(12)	The order of	frate of diffu	sion of gases N.	H_3 , SO_2 , $C\ell_2$	and CO2 is:	- (A) NH ₃ > S	$SO_2 > C\ell_2 > CO_2$
							$O_z > C\ell_\gamma > SO_\gamma$
(13)			water at 1 atm			* * * * * * *	01 - 01 2 - 50 1
	(A) 98°C		(B) 100°C	3	(C) 69°C	(D) 12	10°C
(14)	Ionic solids	are characteri	zed by:- (A) Low meltin		(D) 1.	.0 C
						Solubility in pol	ar solvents
(15)			lete, the enterin			erante un violet von met 19 mm militär (film 1977). T	NACOLING PROPERTY CONTROL
	(A) 7f		(B) 7s		(C) 7p	(D) 7	d
(16)	The velocity	of Photon is:	- (A) Inc	lependent of it	s wavelength	19_30 E.	

(B) Equal to square of its amplitude

(17)

(A) *HCℓ*

____ Hydrogen Halide has the highest percentage of ionic character.

(B) HBr

(C) HF (D) HI
22(Obj) (2017(A)-12000 (MULTAN)

(D) Depends on its source

(C) Depends on its wavelength

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Roll	No.		
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2488 INTERMEDIATE PART-I (11th CLASS)

CHEN	MISTRY	PAPER-I	(NEW SC	HEME)	(SESSION	V 2015-20	17)	GROUP-II
TIME	ALLOWE	D: 20 Minut	es	OBJECT	CIVE	MAX	KIMUM	MARKS: 17
think is Cutting as give	s correct, fill g or filling to n in objectiv LES are not	that circle in vo or more ci e type question	front of that reles will resu on paper and	ve type question question num dt in zero mar leave others bl on on this shee	ber. Use mar k in that ques lank. No cred	ker or pen tion. Atter lit will be	to fill the npt as ma warded i	circles. iny questions
(1)		tion constant o	of pure water a	t 25°C is:-	(A) 1.8 ×	10 ⁻¹⁶ mole	dm^{-3}	
	(B) 1.6 × 1	0 ⁻¹⁴ mole dm ⁻	(C) 1.0	$\times 10^{-14} mole^2$	dm^{-6}	(D) 1.8 ×	10 ⁻¹⁴ mole	² dm ⁻⁶
(2)	The molal	boiling point o	constant is the	ratio of the ele-	vation in boili	ng point to:		
	(A) Molari	ty	(B) Molali	ty (C) Mole fi	raction of solu	te (D) M	ole fractio	on of solvent
(3)	The oxidat	ion number of	Fluorine in O	0F ₂ is:-	(A) +2 (A)	B) -2	(C) -1	(D) + 1
(4)	The unit of	the rate const	ant is the same	e as that of rate	of reaction in	1 - 0		
(5)		of one mole o		er reaction	(C) Zero o	order reacti	on (D) 3 ^r	d order reaction
	(A) 1.008	g	(B) 0.55 m	ig .	(C) 0.184	mg	(D) 1.67	3mg
(6)	The larges	t number of m	olecules are p	resent in:-				
	(A) 3.6 g	of H_2O	(B) 4.8 g (of C ₂ H ₅ OH	(C) 2.8 g	of CO	(D) 5.4 g	g of N_2O_5
(7)	Solvent ex	traction is an	equilibrium pr	ocess and it is	controlled by:-	(A) La	w of Mass	s Action
		nount of Solve		C) Distribution		(D) The an		
(8)	If absolute	temperature o	of a gas is doul	oled and pressu	re is reduced t	o one half,	the volum	ne of gas will:-
	(A) Remai	n unchanged	(B) Increas	e four times	(C) Reduc	e to $\frac{1}{4}$	(D) Be d	oubled
(9)	The order	of rate of diffu	sion of gases	NH_3 , SO_2 , C	ℓ_2 and CO_2	is:- (A) A	$H_3 > SO_2$	$> C\ell_{\gamma} > CO_2$
				$C\ell_2 > SO_2 >$				
(10)	The boilir	ng point of pur	e water at 1at	m pressure is	;-		· · · · · · · · · · · · · · · · · · ·	
	(A) 98°C		(B) 100°C	00 0 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1	(C) 69°C		(D) 120°	C
(11)	Ionic solic	ls are characte	ASSESSMENT OF THE PROPERTY OF THE PARTY OF T	(A) Low me			(0) 120	
				(C) High vapou) Solubility	in polar s	solvents
(12)				ring electron go			a afi	
	(A) 7f		(B) 7s		(C) 7p		(D) 7d	
(13)	The veloci	ty of Photon is	s:- (A)	Independent of	its wavelengt	h		
	(B) Equal to	o square of its	amplitude	(C) Depends o	n its waveleng	gth (D) I	Depends o	n its source
(14)	Hydr	ogen Halide h	as the highest	percentage of i	onic character	•		
	(A) $HC\ell$		(B) HBr		(C) HF		(D) HI	
(15)	The bond or	der of N_2 is:-		(A) 2	(B) 3	(C) 4	(D) 1	
(16)	For the reac	tion NaOH +	$HC\ell \longrightarrow I$	$VaC\ell + H_2O$ t	he change in e	enthalpy is	called:-	
				ation (C) Hea				ombustion
(17)				s solution of H				
	(A) 3.0		(B) 2.7		(C) 2.0		(D) 1.5	

BOARD OF INTERMEDIATE AND SECONDARY EDUCATION,

MULTAN (New Scheme)
OBJECTIVE KEY FOR INTER (PART-1/II) Annual Examination, 2017.

Name of Subject Chemistry Group: 1st Paper Paper Paper Paper Q. Nos. Code Code Code Code 2481 2483 2485 2487

1.	C	B	A	B
2.	D	C	В	C
3.	C	A	В	B

4.	8	B	C	C
5.	C	C	B	C

6.	A	A	C	D
7.	8	D	-	-

10.70	0	B	C.	C
8.	C	D	D	B
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	-	-		
9.	A	A	C	C
10.	B	B	B	A

	13	D	0	
11.	D	B	C	B
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1.771	\Box	(1	C
13.	B	B	B	A
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14.	B	C	C	B
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15,	C	\subset	A	D
16.	B	D	B	A

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18.		
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20.		

1-1/ 11/ Auto	dai madallima
Session	2015-17
Group:	2nd

Q. Nos.	Paper Code	Paper Code	Paper Code	Paper Code
1.155	2482	2484	2486	2488
1.	В	В	В	C
2,	A	В	C	B
3.	C	В	B	B
4.	В	D	c	C
5.	В	C	В	B
6.	B	A	C	А
7.	D	C	C	C
8.	C	В	B	B
9,	A	C	A	В
10.	C	В	C	В
, II.	B	C	В	D
12.	B	В	B	C
13.	В	C	В	A
14.	C	C		A
15.	В	В	D	В

C

B

(Subjective & Objevtive) كويظر عميق چيك كرايات يد پر چيليس كيين مطابق Set كيا كيا ب-اس واليد پرچديس كي من كول غلطی نہے ہم نے سوالیہ پر چیکااردواہ احمریز ک Version بھی چیک کرایا ہے یہ Version آپس مطابقت رکھتے ہیں اور سلبس (Syllabus) ے مطابق بھی ہیں۔ نیزاس پرچے کا Key کیابت بھی تعدیق کا جاتی ہے کہ ریجی درست بنائی گئے ہے۔ اس میں بھی کی کو فی فلطی نہے۔ مریدید کہ ہم نے Key بنا نے سے متعلق وفتر کی جانب سے تیارہ کردہ ہدایات وصول کر کے ان کا بغور مطالعہ کرلیا ہے اور مان کی روشنی میں Kay بنائی ہے۔

16.

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Roll	No:		
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INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (OLD SCHEME)

(SESSION 2012-2014)

GROUP-I

TIME ALLOWED: 3.10 Hours

SUBJECTIVE

MAXIMUM MARKS: 83

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

SECTION-I

2. Attempt any eight parts.

 $8 \times 2 = 16$

- (i) How do you deduce the fractional atomic masses of elements from the relative isotopic abundances? Give example of Ne.
- (ii) Isotopes have same chemical properties. Why?
- (iii) Calculate number of grams in 0.1g of Na.
- (iv) What is Distribution Coefficient? To which techniques it is applied?
- (v) Differentiate between Partition and Adsorption Chromatography.
- (vi) Mention major steps involved in Crystallization. (any four)
- (vii) Verify Boyle's Law using Kinetic equation of Gases.
- (viii) Define Joule Thomson Effect. Give its significance.
- (ix) Write two applications of Plasma.
- (x) Define common ion effect. Give two examples.
- (xi) What is the effect of common ion on solubility of ionic compounds?
- (xii) Why do the rate of forward reactions slow down with passage of time?

Attempt any eight parts.

 $8 \times 2 = 16$

- Give two uses of Liquid Crystals.
- (ii) Why Sodium is softer than Copper? Justify with reason.
- (iii) "Water is liquid at room temperature while H_2S is gas". Explain this statement.
- (iv) Explain cleavage of crystals and cleavage plane.
- (v) Write the nuclear equation for the discovery of Neutrons.
- (vi) What is the Origin of Line Spectrum?
- (vii) Write two properties of Neutrons.
- (viii) Which particles are formed by the decay of free neutron? Give an equation.
- (ix) Define Electrolytic Cell. Give its example.
- (x) Write two functions of Salt Bridge.
- (xi) Calculate the oxidation number of Chromium in $K_2Cr_2O_7$.
- (xii) What are Hydrates? Give example.

4. Attempt any six parts.

 $6 \times 2 = 12$

- Differentiate Bonding Molecular Orbital and Antibonding Molecular Orbital.
- (ii) Why is Sigma Bond stronger than Pi bond?
- (iii) Define Ionization Energy. Why the second ionization energy is higher than that of first?
- (iv) Why the Dipole moment of CO_2 is zero and that of H_2O is 1.85 D?
- (v) Define State and State Function.
- (vi) How will you determine Enthalpy of formation of CO using Hess's Law?
- (vii) Define average rate of reaction. Give its uses.
- (viii) What are Enzymes? Give their two characteristics.
- (ix) How surface area affects the rate of a chemical reaction?

SECTION-II

NOTI	E: - Attempt any three questions of the following:-	
5.(a)	Define the terms with example. (i) Mole (ii) Avogadro's number (iii) Molar Gas Volume (iv) Empirical formula	4
(b)	What are Liquid Crystals? Give their uses. (any three)	4
6.(a)	What are Quantum Numbers? Describe Principle Quantum Number 'n'	4
(b)	250 cm ³ of the sample of Hydrogen effuses four times rapidly as 250 cm ³ of an unknown gas. Calculate the molar mass of an unknown gas.	4
7.(a)	Describe Molecular Orbital Theory to show formation of bonding and antibonding molecular or Which types of molecular orbitals are formed by linear and parallel overlaps of two p – orbitals	bital
(b)	Explain working of bomb calorimeter to estimate heat of combustion of organic compound.	4
8.(a)	The solubility product of Ag_2CrO_4 is 2.6×10^{-2} at $25^{\circ}C$. Calculate the solubility of the compound.	4
(b)	How rate of a chemical reaction depends upon nature of reactants and surface area?	4
9.(a)	What is Boiling Point Elevation? How boiling point elevation of a solution is determined?	4
(b)	Discuss the construction and working of Electrolytic Cell.	4
	SPOTION HI (PRACTICAL PART)	
- 0	SECTION-III (PRACTICAL PART) NOTE: (2) Attempt any three parts (3 x 5 = 1)	5)
10.	NOTE:- (i) Attempt any three parts. (ii) Write down material required, diagram and procedure for part A & B.	
	(iii) Write down standard solution, chemical equation with mole ratio, indicator with end point, procedure and supposed readings with calculations for part C, D & E. (1+1+1+1+1)	= 5
(A	The given solution contains 8 g of Oxalic Acid and Sodium Oxalate in 1000 Cm ³ . Determine the percentage of Sodium Oxalate.	
(E	Determine the value of x in a solution containing 27.8 g of $FeSO_4.xH_2O$ dissolved per G	lm³.
(0	Determine the amount of I_2 per dm^3 in the given sample solution. You are provided with 0. $Na_2S_2O_3$ solution.	l M
(I	D) Purify a given sample of common salt by passing HCℓ gas:-	
а	 Separate and identify lead and cadmium ions in a mixture of their salts by paper chromatogra 	phy.

21-2017(A)-3200 (MULTAN)

Panar	Code 2017 (A) Roll No.
,	(AO1 INTERMEDIATE PART-I (11th CLASS)
Numb	per: 0401
ΓΙΜΕ Note: think i: Cuttin;	MISTRY PAPER-I (OLD SCHEME) (SESSION 2012-2014) GROUP-I ALLOWED: 20 Minutes OBJECTIVE MAXIMUM MARKS: 1 You have four choices for each objective type question as A, B, C and D. The choice which you seement, fill that circle in front of that question number. Use marker or pen to fill the circles. If you go refilling two or more circles will result in zero mark in that question. Attempt as many questions in objective type question paper and leave others blank. No credit will be awarded in case
BUBBI	LES are not filled. Do not solve question on this sheet of OBJECTIVE PAPER.
Q.No.1 (1)	27 g of $A\ell$ will react completely with mass of O_2 to produce $A\ell_2O_3$.
(-)	(A) 8 g of Oxygen (B) 16 g of Oxygen (C) 24 g of Oxygen (D) 32 g of Oxygen
(2)	The comparative rates at which the solutes move in paper chromatography mainly depends on:-
(4)	
	Carried Newson Contraction Con
(2)	(C) Temperature of the experiment (D) Size of chromatographic tank used
(3)	The partial pressure of Oxygen in the lungs is:- (A) 760 torr (B) 310 torr (C) 159 torr (D) 116 torr
(4)	(A) 760 torr (B) 310 torr (C) 159 torr (D) 116 torr Vapour pressure of a substance does not depend upon:-
(.)	(A)Temperature (B) Inter-molecular forces (C) Surface area (D) Geometrical shape of molecules
(5)	The quantum number values for $2p$ orbitals are:-
N-2	(A) $n = 2$, $\ell = 1$ (B) $n = 1$, $\ell = 2$ (C) $n = 1$, $\ell = 0$ (D) $n = 1$, $\ell = 1$
(6)	
(0)	Hydrogen halides has the highest percentage of ionic character. (A) IIF (B) HC\ell (C) HBr (D) HI
(7)	(A) IIF (B) HCℓ (C) HBr (D) HI If an endothermic reaction is allowed to take place very rapidly in the air, the temperature of the
6.7	surrounding air:- (A) Remains constant (B) Increases (C) Decreases (D) Remains unchanged
(8)	pH of 0.001 M NaOH solution is:- (A) 3 (B) 1 (C) 11 (D) 12
(9)	The molal boiling point constant is the ratio of the elevation in boiling point to:-
200	(A) Molarity (B) Molality (C) Mole fraction of solute (D) Mole fraction of solvent
(10)	Stronger the oxidizing agent, greater is the:-
	(A) Oxidation potential (B) Reduction potential (C) Redox potential (D) EMF of the cell
(11)	The unit of rate constant is same as that of the rate of reaction in:-
	(A) Zero order reaction (B) First order reaction (C) Second order reaction (D) Third order reaction
(12)	The number of $C\ell^-$ ions per unit cell in $NaC\ell$ is:- (A) 2 (B) 4 (C) 6 (D) 8
(13)	One mole of SO_2 contains:- (A) 6.02×10^{23} atoms of Oxygen
	(B) 18.1×10^{23} molecules of SO_2 (C) 6.02×10^{23} atoms of Sulphur (D) 4 gram atoms of SO_2
(14)	Number of molecules in one dm^3 of water is close to:-
138 16	(A) $\frac{6.02 \times 10^{23}}{22.4}$ (B) $\frac{12.04}{22.4} \times 10^{23}$ (C) $\frac{18}{22.4} \times 10^{23}$ (D) $55.6 \times 6.02 \times 10^{23}$
(15)	The nature of the positive rays depend upon the:- (A) Nature of the electrode
No. T.	(B) Nature of the discharge tube (C) Nature of the residual gas (D) All of these
(16)	Each Carbon atom of Ethene has Hybridization:-
ar #0.000	(A) sp (B) sp^2 (C) sp^3 (D) dsp^2
(17)	A solution of pH = 0 indicates molar concentration of H^+ ions:
(.,)	THE NAME OF THE PARTY OF THE PA
	(2)
	21(Obj)(\(\frac{1}{12}\))-2017(A)-3200 (MULTAN)

V							
Paper	Code			2017 (A)		Roll No.	
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Numb	er;	0403			-57. N. CO. 11. 1. A. S.		
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Q.No.1							
(1)			positive rays depend	- F	707 700	Nature of the electron	ode
100			discharge tube (C)		residual gas	(D) All of these	
(2)	Each	Carbon aton	n of Ethene has Hybri	idization:-			
	(A)	sp	(B) sp^2	(C) sp3		(D) dsp ²	
(3)	A sol	lution of pl	I = 0 indicates mola	r concentration	of H ion	ns:-	
	(A)	10^{-7}	(B) 10 ⁷	(C) 1.0		(D) 10 ⁻¹⁴	8
(4)	27 g	of Al will	react completely with	h mass o	of O, to pr	oduce $A\ell_1O_1$.	
(5)	(A)	8 g of Oxyge	en (B) 16 g of Oxygerates at which the sol	en (C) 24 g o	of Oxygen	(D) 32 g of Oxyge	
1.0		Size of paper					epends on:-
	(0) (0) (2)(e)(0)	18 8). 20			lues of solu		
Tare.			of the experiment		f chromatos	graphic tank used	
(6)			are of Oxygen in the l	191			
7 7 0	1000	60 torr	(B) 310 torr	(C) 159 to		(D) 116 torr	
(7)	0.8888870		of a substance does no	10 To			
(8)	AT MOST 185	emperature Juantum nun	(B) Inter-molecular their values for $2p$ or		ırface area	(D) Geometrical s	hape of molecules
	(A) n	$=2$, $\ell=1$	(B) $n = 1$, $\ell = 2$	(C) $n = 1$,	$\ell = 0$	(D) $n = 1$, $\ell = 1$	
(9)		Hydrogen h	alides has the highest	percentage of	ionic chara	cter.	
	(A) H		(B) <i>HCℓ</i>	(C) HBr		(D) HI	
(10)	If an	endothermic	reaction is allowed to	o take place ve	ry rapidly i	n the air, the temper	rature of the
		unding air:-	(A) Remains cons				
(11)	pH o	f 0.001 M	NaOII solution is:-	(A) 3		C) 11 (D) 12	
(12)	The n	nolal boiling	point constant is the	ratio of the ele	vation in bo	oiling point to:-	
		olarity	(B) Molality			olute (D) Mole fra	action of solvent
(13)	Stron	ger the oxidi	zing agent, greater is				
	(A) ()	xidation pot	ential (B) Reduction	n potential (C) Redox po	tential (D) EMF o	of the cell
(14)			onstant is same as tha				
	(A) Z	ero order re	action (B) First ord	er reaction (C) Second or	der reaction (D) I	hird order reaction
(15)	The n	umber of C	ℓ^- ions per unit cell i	n <i>NaCℓ</i> is:-	(A)	2 (B) 4 (C)	6 (D) 8
(16)	One n	nole of SO2	contains:-	(A) 6.02 × 1	0 ²³ atoms	of Oxygen	
	(B) 18	$1 \times 10^{23} \text{ mg}$	elecules of SO ₂ (C)	6.02 × 10 ²³ at	toms of Sul	phur (D) 4 gran	a atoms of SO.
			ules in one dm ³ of wa				
			(B) $\frac{12.04}{22.4} \times 10^{23}$			(D) 55.6 × 6.02 × 1	025

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Paper	Code				20)17 (A)		R	oll No.	
Numb	er:	648	5	INTERN			T-I (11 th C			
TIME Note: think is Cutting as give	ALLO You be s corre g or fil n in ob LES ar	OWED: 2 nave four o ect, fill tha ling two o ojective ty	20 Mi choice at circl or mor pe que	s for each e in front e circles w estion pap	objective of that q vill result er and le	OBJE e type qu uestion n in zero n ave other	CTIVE estion as A, l umber. Use	B, C and D marker or question. A credit will	MAXIMU The choice pen to fill Attempt as be awarde	many questions
(1)	If an	endothern	nic rea	ection is al	lowed to	take place	very rapidly	in the air, t	he temperat	ure of the
		ounding air	POS P			nt (B) In	creases (C)	Decreases	(D) Remai	ns unchanged
(2)				H solution		(A) 3	(C) 20	201200	(D) 12	
(3)						tio of the	elevation in	ooiling poin	t to:-	
9885		Aolarity		(B) Molali	(5)		le fraction of	solute (D)	Mole fract	ion of solvent
(4)				g agent, gr						
(5)	The	unit of rate	const	ant is sam	e as that o	f the rate	(C) Redox p of reaction in (C) Second	1:-		the cell ird order reaction
(6)				ions per u					4 (C) 6	
(7)		mole of S					1×10^{23} atom	XX100000000000000000000000000000000000	Contract of the Contract of th	(D) 6
			***					Action in the second	W.)	atoms of SO,
(8)				in one dn				apilar (i) 4 gram a	tions of SO ₂
(9)		H151100 1001124	3							
(8)							$\frac{3}{4}\times10^{23}$			
(9)							(A)			
27 W							e residual gas	(D) All	of these	
(10)				Ethene ha						
	(A)			B) <i>sp</i> ²		(C) <i>sp</i>		(D) <i>dsp</i> ²		
(11)	A so	lution of	pH =	indicate	s molar c	oncentrat	ion of H+ io	ons:-		
	(A)	10-7	(B) 10 ⁷		(C) 1.0		(D) 10 ⁻¹⁴		
(12)	27 g	of Al wi	ill reac	t complete	ely with _	mas	s of O2 to p	oduce $A\ell_2$	O_3 .	
	(A) 8	g of Oxy	gen (B) 16 g of	Oxygen	(C) 24 g	g of Oxygen	(D) 32 g o	of Oxygen	
(13)							paper chrom			ends on:-
		ize of pape					values of solu		3 6	
	(C) T	emperatur	e of th	e experime	ent	(D) Size	of chromato	graphic tan	k used	
(14)	The p	artial pres	sure o	f Oxygen i	n the lung					
	(A) 7	60 torr	(1	3) 310 torr		(C) 159	torr	(D) 116 to	rr	
(15)	Vapo	ur pressure	e of a s	substance o	loes not d	lepend up	on:-	***************************************		
								(D) Geom	etrical shap	e of molecules
(16)				values for					•	
	(A) n	= 2, <i>l</i> = 1	(F	3) $n = 1$, ℓ	= 2	(C) n =	1, $\ell=0$	(D) $n = 1$,	$\ell = 1$	
(17)		Hydrogen	halide	s has the h	ighest per	rcentage o	of ionic chara	cter.		
	(A) H	F	(H	B) <i>HCℓ</i>		(C) HB		(D) HI		

Paper	Code	20	017 (A)	Roll No.
Numbe	er: 6487		TE PART-I (11th C	
		EDI (OIDSCH	EME) (SESS	ION 2012-2014) GROUP-I
	ALLOWED: 20 M		OBJECTIVE	MAXIMUM MARKS: 1
			e type question as A,	B, C and D. The choice which you
				marker or pen to fill the circles. question. Attempt as many questions
as given	in objective type q	uestion paper and le	eave others blank. No	credit will be awarded in case
Q.No.1	are not illed.	Do not solve question	n on this sheet of OB.	JECTIVE PAPER.
(1)	The quantum numb	per values for 2p orb	itals are:-	
	(A) $n = 2$, $\ell = 1$	(B) $n = 1$, $\ell = 2$	(C) $n=1$, $\ell=0$	(D) $n = 1$, $\ell = 1$
(2)	Hydrogen hal	lides has the highest p	ercentage of ionic char	racter.
	(A) HF	(B) <i>HCℓ</i>	(C) HBr	(D) HI
(3)	If an endothermic	reaction is allowed to	take place very rapidly	in the air, the temperature of the
946				Decreases (D) Remains unchanged
(4)		nOH solution is:-		(C) 11 (D) 12
(5)			atio of the elevation in	
(6)	NO. 1	(B) Molality		f solute (D) Mole fraction of solvent
(6)		ing agent, greater is the		= A [A A/] - 1205 B05550 NOM - 7000
(7)				potential (D) EMF of the cell
(1)			of the rate of reaction i	
(8)				order reaction (D) Third order reaction
		ions per unit cell in		A) 2 (B) 4 (C) 6 (D) 8
(9)			(A) 6.02×10^{23} atom	Carried Co. C. C. C. C.
				ulphur (D) 4 gram atoms of SO ₂
(10)	Number of molecul	les in one dm ³ of wat	er is close to:-	
	(A) $\frac{6.02 \times 10^{23}}{22.4}$	(B) $\frac{12.04}{22.4} \times 10^{23}$	(C) $\frac{18}{22.4} \times 10^{23}$	(D) $55.6 \times 6.02 \times 10^{25}$
(11)	The nature of the pe	ositive rays depend up	oon the:- (A)	Nature of the electrode
	(B) Nature of the di	ischarge tube (C) Na	ature of the residual ga	s (D) All of these
(12)	Each Carbon atom	of Ethene has Hybridi	zation:-	
36	(A) sp	(B) sp ²	(C) sp ³	(D) dsp ²
(13)	A solution of pH	= 0 indicates molar c	concentration of H^+ is	ons:-
	(A) 10 ⁻⁷	(B) 10 ⁷	(C) 1.0	(D) 10 ⁻¹⁴
(14)	27 g of Al will re	act completely with _	mass of O_2 to p	roduce $A\ell_2O_3$.
			(C) 24 g of Oxygen	
(15)				matography mainly depends on:-
	(A) Size of paper us		(B) R_f values of solu	
	(C) Temperature of	the experiment	50 W W	
		of Oxygen in the lun		
	(A) 760 torr	(B) 310 torr	(C) 159 torr	(D) 116 torr

(17) Vapour pressure of a substance does not depend upon:-

Roll No:

INTERMEDIATE PART-I (11th CLASS)

(SESSION 2012-2014) GROUP-II PAPER-I (OLD SCHEME) CHEMISTRY

TIME ALLOWED: 3.10 Hours

SUBJECTIVE

MAXIMUM MARKS: 83

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

SECTION-I

2. Attempt any eight parts. $8 \times 2 = 16$

- (i) Why actual yield of a chemical reaction may be less than the theoretical yield?
- (ii) What is a Molecular Ion? How is it generated?
- (iii) Define Stoichiometry. What are assumptions necessary for Stoichiometric Calculations?
- What is a good method for drying of the crystallized substance? (iv)
- (v) Describe briefly, the method of sublimation.
- (vi) Name any four methods for separation and purification of products.
- (vii) Write quantitative definition of Charles's law.
- (viii) Derive SI units of General Gas Constant R.
- (ix) Write down formulas for converting ${}^{o}C$ into F^{o} and F^{o} into centigrade.
- (x) What is the effect of change of pressure on the direction of following reaction? $2SO_2 + O_2 \implies 2SO_3$
- Define pH and pOH. (xi)
- (xii) What is meant by Ionization Constant of Acids?

3. Attempt any eight parts.

 $8 \times 2 = 16$

- (i) Evaporation of a liquid cause cooling. Explain it.
- (ii) What are dipole - induced dipole forces?
- Define Transition Temperature. Give an example. (iii)
- (iv) Cleavage of crystals is itself anisotropic behaviour. Explain it.
- (v) What is Pauli Exclusion Principle?
- Write two defects of Rutherford Atomic Model. (vi)
- (vii) What is Stork Effect?
- (viii) Define Frequency and Wave number.
- Sum of mole fraction is equal to unity. Justify the statement. (ix)
- What is Discontinuous Solubility Curve? Give an example. (x)
- Define Oxidation Number. What is oxidation number of Phosphorous in H_3PO_4 . (xi)
- Write the reactions of electrodes of silver oxide battery. (xii)

4. Attempt any six parts.

 $6 \times 2 = 12$

- (i) Why the Electron Affinity of Chlorine is greater than Fluorine?
- (ii) The dipole moment of CO_2 and CS_2 are zero, but that of SO_2 is 1.61 D. Give reason.
- (iii) Define Electronegativity. How it varies along the period of periodic table?
- (iv) Write Lewis structure of
 - (a) H_2SO_4 (b) N_2O_5
- Justify that heat of formation of a compound is the sum of all the other enthalpies. (v)
- (vi) How will you differentiate between ΔE and ΔH ?
- (vii) Differentiate between Homogeneous and Heterogeneous Catalysis.
- Rate of chemical reaction is an ever changing parameter under the given condition. (viii) Justify the statement.
- (ix) What do you mean by Autocatalysis? Explain with an example.

SECTION-II NOTE: - Attempt any three questions of the following:-What is Limiting Reactant? How does it control the quantity of the product formed? 5.(a) 4 Explain with example. 4 What are Liquid Crystals? Give their uses in daily life. (b) Calculate the density of CH_4 at $0^{\circ}C$ and 1 atmospheric pressure. What will happen to the 6.(a) 4 density if temperature is increased to 27°C? (b) What are Positive Rays? How are they produced? Write down their properties. 4 7.(a) Draw molecular orbital diagram showing relative energies of O_2 and N_2 . Describe measurement of Enthalpy of reaction by Glass Calorimeter. (b) Solubility of CaF_2 in water at $25^{\circ}C$ is found to be 2.05×10^{-4} mol dm^{-3} . 8.(a) What is value of K_{sp} at this temperature? Write a brief note on energy of Activation. (b) What is Molal Boiling Point Constant? How can we determine the Molar Mass of 9.(a) 4 an unknown solid by elevation of Boiling Point Method? 4 What is Electrochemical Series? Give its three applications. (b) SECTION-III (PRACTICAL PART) $(3 \times 5 = 15)$ NOTE:- (i) Attempt any three parts. 10. (ii) Write down material required, diagram and procedure (1+1+3)=5for part A & B. (iii) Write down standard solution, chemical equation with mole ratio, indicator with end point, procedure and supposed readings with (1+1+1+1+1)=5calculations for part C, D & E. (A) Prepare Crystals of Benzoic Acid from Saturated Solution. Separate a mixture of inks by Paper Chromatography. (C) The given solution contains 10.0 g of washing soda dissolved per dm³. Determine % of Na, CO3 in the given sample. (D) The given solution contain 30.0 g of FeSO₄. 7H₂O per dm³. Find out % age of Fe in the given sample.

Determine % age purity of the sample.

(E) The given solution contains 30.0 g of Na₂S₂O₃.5H₂O dissolved per dm³.

Paper			FATOR		2017 (A)	no a caeth co		Roll No		
Numbe	er:	6482	2 INT	ERMEDIA	TE PAR	T-I (11 th Cl	LASS)			
CHEM	IISTR	Y PA	PER-I	OLD SCH	IEME)	(SESSIC	ON 20	12-2014)	GRO	UP-II
2			0 Minutes		OBJEC			MAXIMU		
						estion as A, B umber. Use r				
Cutting	or filli	ng two or	more circ	les will resu	lt in zero n	nark in that q	uestion	. Attempt as	many q	uestions
as given BUBBL	ES are	ective typ not filled	e question	paper and solve questic	leave other on on this s	s blank. No cheet of OBJE	credit w	vill be awarde E PAPER.	ed in cas	ie
Q.No.1 (1)		es differ i								
(1)	- some i			nd unon mas	o (D) A	angement of e	lantene	. i		
					. 8 2	. 5				2.13
(2)						h they may be	апесте	d in electrom	agnetic t	ield
(2)			upled by 1	.4 g of N ₂ a						ts:
(2)	(A) 2.					(C) 1.12 dm ³		(D) 112 cm ³		
(3)						is controlled b	y;-	(A) Law o	f mass a	ction
(4)	M70000 1000		of solvent	A CONTRACTOR OF THE PARTY OF TH	C) Distribu			(D) The amo		lute
(4)						the volume of		vill become tv	vice of	
						(C) 546 I		(D) 273 K		
(5)	The or	der of the	rate of diff	fusion of gas	es NH_3 , S	O_2 , $C\ell_2$ and	CO ₂ is	S:-		
	(A) N	$H_3 > SO_2$	$> C\ell_2 > 0$	CO_2	(B) NH	$_3 > CO_2 > SC$	$O_2 > C\ell$	2		
	(C) Ci	$c_2 > SO_2$	$> CO_2 > N$	VH_3	(D) Cl	> CO2 > SO	$rac{1}{2} > NE$	I_3		
(6)	When	water free	zes at 0°C	, its density	decreases d	ue to:-	(A) Cu	hic structure o	ofice	
						Change of bo				ond angle
(7)				dry ice form						
	SAMOONNES SA	ic crystals		valent crysta		lolecular cryst	als ()	D) Any type o	of crystal	e e
(8)	In the g	ground sta		om, the electr	W.1- (20.000 m2)		Total Co	-, and offer	i vij siai	
					200	earest to the n	ucleus	(D) Farthest	from the	e nucleus
(9)				are called:-						
	(A) Hy	orid orbita	ds (B) Va	alence orbita	s (C) Deg	enerate orbital	s	(D) <i>d</i> – orbita	als	
(10)	s	pecies has	s unpaired	electrons in a	ntibonding	molecular orl	oitals.	M. Macro Terranac		
			(B) N		(C) B ₂			(D) F ₂		
(11)	In a gr	oup of per	riodic table	, ionization	energy:-			1000		
	(A) De	ecreases	(B) Inc	reases	(C) Rem	ains same	(D) I	First increases	then dec	creases
(12)	If an en	dothermi	c reaction i	s allowed to	take place	very rapidly in				
		nding air:				reases (C) De				nged
(13)	When	HCℓ is a		S aqueous se						
	(A) Dec	creases	(B) Inc	reases	(C) Rem	ains constant	(D) F	irst increases	then dec	reases
(14)	The pH	of 10 ⁻³ r	nole dm ⁻³	of an aqueou	s solution o	of H_2SO_4 is:-		.0 (B) 2.0		
				nol in water				qual to that of		New York
				(C) More				(D) Less than		vater
) Zinc (B) C		(C) Lead	(D) Gra	
						lyst is called:-		25 : 14 to 12 to 12 to 14 to 15 to 16 to 1		sacriculaes
	(A) Pro	moter	(B) Act	ivator	(C) Inhib	itor		(D) Positive c	atalys:	

8											
Paper	Code				2	017 (A)			Roll No		_
Numbe	er:	64	84	INTE	RMEDIA	TE PA	RT-I (11	" CLASS	S)		
TIME A Note: think is Cutting as given	ALL(You h corre or fill in ob	OWEI nave for ect, fill ling two jective	D: 20 N ur choi that cir to or me type q	Minutes ces for ea cle in fro ore circle uestion	ach objection ont of that es will resul paper and l	OBJE ve type q question It in zero eave other	uestion as number. mark in t ers blank.	A, B, C ar Use marke hat questic No credit	nd D. The cho er or pen to fil	s many question	17
(1)	Ana	aqueou	s soluti	on of Eth	anol in wate	er has vap	our pressu	re:- (/	A) Equal to tha	t of water	
	(B) I	Equal t	that of	Ethanol	(C) Mor	e than tha	at of water		(D) Less tha	an that of water	
(2)	In si	lver ox	ide batt	ery, the a	node is mad	le of:-	(A) Zinc	(B) Copp	er (C) Lead	(D) Graphite	
(3)	The	substar	ice which	h decrea	ses the activ	ity of a c	atalyst is c	alled:-			
	(A)	Promo	ter	(B) Act	tivator	(C) In	hibitor		(D) Positive	catalyst	
(4)			iffer in:		240	02250 0		0 02020 0	20 000000		
					27/12/2014	16.00			ons in orbitals		
(5)								ay be affec	cted in electron	nagnetic field	
(5)					g of N_2 a			22			
	(A)	2.24 dr	n's		(B) 22.4 dn	n^2	(C) 1.12	dm ³	(D) 112cm ³		
(6)					llibrium pro			and the second second	1 12 1-04 1 270 270	of mass action	
					ised (IDTEST A CHARLEST AND I	ount of solute	
(7)									s will become	twice of	
	what	t it is a	$0^{\circ}C$.	(A) :	546° C	(B) 200°	C (C)	546 K	(D) 273 K		
(8)	The	order o	f the rat	e of diffi	ision of gas	es NH_3 ,	SO_2 , $C\ell_2$	and CO2	is:-		
	(A)	$NH_3 >$	SO ₂ >	$C\ell_2 > C$	CO_2	(B) A	$IH_3 > CO_2$	> SO ₂ >	$C\ell_2$		
	(C)	$C\ell_2 >$	SO ₂ > 0	$CO_2 > N$	H_3	(D) C	$C\ell_2 > CO_2$	> SO ₂ > 1	NH_3		
(9)	When	n water	freezes	at 0°C,	its density of	decreases	due to:-	(A) C	ubic structure	of ice	
										ange of bond ang	les
(10)	Charles I Control of the				y ice form t						
	(A) I	onic cr	ystals	(B) Co	valent crysta	als (C)	Molecular	r crystals	(D) Any type	of crystals	
(11)					m, the elect						
	(A) 1	In the r	ucleus	(B) In t	he second s	hell (C)	Nearest to	the nuclei	us (D) Farthe	st from the nucle	us
(12)					are called;						
	(A) I	Iybrid	orbitals	(B) Va	lence orbita	ıls (C) D	egenerate	orbitals	(D) d - orb	itals	
(13)		speci	es has u	npaired e	electrons in	antibondi	ng molecu	lar orbitals	s.		
	(A)	O_2^{+2}		(B) N	-2 1	(C) B	2		(D) F ₂		
(14)	Inag	group (of period	lic table,	ionization e	energy:-					
	(A)	Decrea	ises	(B) Inc	reases	(C) R	emains san	ne (D) First increase	es then decreases	
(15)									The temperate uses (D) Rema	ure of ins unchanged	
(16)					S aqueous so						
	(A) I	Decrea	ses	(B) Inc	reases	(C) R	emains cor	nstant (D) First increase	s then decreases	
(17)	The	pH of	10 ⁻³ mo	le dm ⁻³	of an aoueo	us solutio	on of HS	O is- (A)30 (B)20	(C) 15 (D) 2	7

Paper (Code		******	2017 (A)	m r crath er	Roll No	•
Numbe	er: 64	86	INTERMED	IATE PAR	T-I (11" CL	ASS)	
TIME A Note: Y think is Cutting as given BUBBL	ALLOWED You have fou correct, fill to or filling two in objective	o: 20 M ir choice that circ o or moi type qu	es for each obje- le in front of th	OBJEC ctive type que at question n sult in zero n d leave other	CTIVE estion as A, B, 6 umber. Use manark in that que s blank. No cr	C and D. The carker or pen to estion. Attempedit will be award	MUM MARKS: 17 choice which you fill the circles. t as many questions arded in case
Q.No.1 (1)	The molecu	les of Co	O ₂ in dry ice for	m the:-			
	(A) Ionic cry		(B) Covalent cry		Molecular crysta	ds (D) Any ty	pe of crystals
(2)	In the groun		f an atom, the el				
13 0	(A) In the n	ucleus	(B) In the secon	d shell (C)	Nearest to the nu	icleus (D) Far	thest from the nucleus
(3)			energy are calle				CONTRACTOR OF CO
		1025	(B) Valence orl		generate orbitals	(D) d - c	orbitals
(4)			npaired electrons		ACC ACC ACC ACC		
	(Λ) O ₂ ⁺²		(B) N ₂ ⁻²	(C) B ₂		(D) F ₂	
(5)	In a group	of period	lic table, ionizati	on energy:-			
8.0	(A) Decrea	100	(B) Increases	7217.65	nains same	(D) First incre	eases then decreases
(6)		ermic re	action is allowed	d to take place	very rapidly in	air. The tempe	
(7)	When HCl	is added	to H ₂ S aqueous	s solution, its	ionization:-		
	(A) Decreas					(D) First incre	ases then decreases
(8)	The pH of 1	0 ⁻³ mol		andomes:		No. of the second second	2.0 (C) 1.5 (D) 2.7
(9)			n of Ethanol in w			- COUNTY	
0.3			Ethanol (C) N	355	1977	N N N	than that of water
(10)			ry, the anode is t			100000000	
(11)			decreases the a		Harrier Control Manager		***
(12)	(A) Promot	ter	(B) Activator	(C) Inh			tive catalyst
(12)	COLOR DOSE		h depend upon r	nass (B) Ari	angement of ele	ectrons in orbita	ls
				season services			romagnetic field
(13)			ed by 1.4 g of 1		- 24 P		
\/ 0.004556	(A) 2.24 dn				(C) 1.12 dm ³	(D) 112	cm ³
(14)			s an equilibrium			0.00	
(17)			olvent used	000000000000000000000000000000000000000	Ted Carl	000000000000000000000000000000000000000	aw of mass action amount of solute
(15)	9,5		constant, at which	20.00		9 0	
(13)			(A) 546" C				
(16)						Let the second and th	
(16)			of diffusion of				
	(A) $NH_3 >$				$H_3 > CO_2 > SO$	8 8	
	(C) $C\ell_2 >$	$SO_2 > C$	$O_2 > NH_3$	(D) C	$_2 > CO_2 > SO_2$	$> NH_{i}$	
(17)	When water	freezes	at 0°C, its dens	ity decreases	due to:- (A) Cubic struct	ure of ice

(B) Empty space present in the structure of ice (C) Change of bond lengths (D) Change of bond angles 22(Obj)(ななな)-2017(A)- 560

Paper	Code		2017 (A)	ak.	Roll No		
Numbe	er: 6488	INTERMEDI	ATE PART	-I (11 th CLA	ASS)		
CHEM	IISTRY PAPE	R-I (OLD SC	HEME)	(SESSION	N 2012-2014)	GROUP-II	
	ALLOWED: 20 M		OBJEC'			JM MARKS: 17	
Note:	You have four choice	es for each object	tive type ques	tion as A, B, C	and D. The cho	ice which you	
	correct, fill that cir or filling two or me						
as given	in objective type q	uestion paper and	leave others	blank. No cre	dit will be award	ed in case	
Q.No.1	ES are not filled.	Jo not solve quest	ion on this sh	eet of OBJEC	IIVE PAPER.		
(1)	Pressure remaining	constant, at which	temperature t	ne volume of a	gas will become t	wice of	
	what it is at $0^{\circ}C$.	(A) 546°C	(B) $200^{\circ}C$	(C) 546 K	(D) 273 K		
(2)	The order of the rat	e of diffusion of ga	uses NH ₃ , SC	Q_2 , $C\ell_2$ and $C\ell_2$	CO ₂ is:-		
	(A) $NH_3 > SO_2 >$	$C\ell_2 > CO_2$	(B) NH ₃	> CO ₂ > SO ₂	> C\(\ell_2\)	¥.	
	(C) $C\ell_2 > SO_2 > C$	$CO_2 > NH_3$	(D) Cl,	$> CO_2 > SO_2$	> NH ₃		
(3)	When water freezes	- 15	10.00 0E)		A) Cubic structure	ofice	
(2)	(B) Empty space pre			December 1970			
(4)	The molecules of C			ondinge or bone	a longuis (b) on	ange of bothe angre	
36160	(A) Ionic crystals	- Marking and - Artin		decular ervetul	c (D) Any type	of arretals	
(5)	In the ground state				s (D) Any type	of crystals	
(-)	(A) In the nucleus				cleus (D) Farthe	st from the nucleus	
(6)	Orbitals having sam			arost to the m	eleds (D) Farme	se non the necreus	
W.C.	(A) Hybrid orbitals			nerate orbitals	(D) <i>d</i> – orb	itals	
(7)		paired electrons in	10 2011 S45.5		404		
		(B) N ₂ ⁻²	(C) B,		(D) F,		
(8)	In a group of period	ic table, ionization	05/05 ±5		\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \		
	(A) Decreases	(B) Increases	200	ins same	(D) First increase	s then decreases	
(9)	If an endothermic rea	19 COMO I COOKING PROPERTY					
	surrounding air:-				The Control of the Co		
(10)	When HCℓ is add				7.42		
	(A) Decreases	(B) Increases	(C) Rem	ins constant	(D) First increase	s then decreases	
(11)	The pH of 10 ⁻³ mo	le dm^{-3} of an agus				(C) 1.5 (D) 2.	
(12)						0 10 10 M	
4. October	An acueous solution of Ethanol in water has vapour pressure:- (B) Equal to that of Ethanol (C) More than that of water (D) Less than that of water						
(13)	In silver exide batte	and the control of th				(D) Graphite	
(14)	The substance whic				FF (=/=	(2) Grapinie	
	(A) Promoter	(B) Activator	(C) Inhib		(D) Positive	catalyst	
(15)	Isotopes differ in:-				10 8 (11 8 (11) 12 (12) 13 (13)	00.000 0 1 0.000	
	(A) Properties which	ch depend upon ma	ass (B) Arran	gement of elec	trons in orbitals		
	(C) Chemical prop	erties (D) The ex	xtent to which	they may be as	ffected in electrom	agnetic field	
(16)	The volume occupi					120	
	(A) 2.24 dm ³	(B) 22.4 a	lm^3 (6	(2) 1.12 dm ³	(D) 112 cm ³		

(17) Solvent extraction is an equilibrium process and it is controlled by:(B) The amount of solvent used
(C) Distribution law
(D) The amount of solute

(MULTAN)

BOARD OF INTERMEDIATE AND SECONDARY EDUCATION,

MULTAN (Old Scheme) OBJECTIVE KEY FOR INTER (PART-I/M) Annual Examination, 2017,

Groi Q. Nos.	Paper Code	Code	Paper Code	Code
	6481	6483	6485	6487
1.	C	C	C	A
2.	B	B	C	A
3.	\mathcal{D}	C	B	AC
1.	$\widetilde{\mathcal{D}}$	C	B	C
5.	A	B	A	B
).	A	\mathcal{D}	B	B
7,	C	\mathcal{D}	C	A
	C	A	\mathcal{D}	B
	B	A	C	C
0.	B	С	B	\mathcal{D}
1.	A	C	C	C
2.	B	B	C	B
3.	C	B	B	С
1.	D C	A	\mathcal{D}	C
5.		B	\mathcal{D}	B
5,	B	C	A	D
	_	D	A	D

Group Q.	Paper	Paper	Paper	Paper
Nos.	Code	Code	Code	Code
	6482	6484	6486	6488
1,	A	C	C	C
2.		A	C	B
3.	C	A	C	B
4.	C	A	B	C
5.	B	ACCC	A	CCC
6.	B C C	C	C	C
7.	C	C	CAAOU	B
8.	C	B	D	A
9,	C	B	C	C
10.	B	C	A	CA
11.	A	C	AC	D
12.	C	C	A	DC
13.	A	B	C	
14.	D	AC	A C C	A C A C
15.	C	C	C	A
16.	A	A	B	C
17.	C	A D	B	C
18.			,0	_
19.				
26.				

Session (2012-14)

سرفیفیکیٹ بابت تھیج سوالیہ برجہ مارکنگ Key

يم ي منون hemila ي ي المار من المار المارك (Subjective & Objevtive) کو بنظر میں چیک کرلیا ہے یہ پر جیسلیس کے مین مطابق Set کیا گیا ہے۔اس سوالیہ پر چہ می کی حم کی کوئی غطی نہے ہم نے موالید پرچیکا اردوادر اگریز Version بھی چیک کرلیا ہے یہ Version آئیں میں مطابقت دکھتے ہیں اور سلیس (Syllabus) ے مطابق بھی ہیں۔ نیزاس پر چدکی Key کی بایت بھی تقد بین کی جاتی ہے کہ یہ بھی درست بنائی گئی ہے۔ اس میں بھی کی کو کی فلطی نسب-مزید بیکہ ہمنے Key بنانے سے متعلق دفتر کی جانب سے تیارہ کردہ ہدایات وصول کر کے ان کا بغور مطالعہ کرلیا ہے اوران کی روشنی میں Key بنائی ہے۔

PREPARED &	CHECKED	BY
C 41 C.		

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